



**HYDROBORATION, A BRIEF
HISTORICAL REVIEW THROUGH
MECHANISTIC VIEWS, PART I:
ALKYL- AND ARYL-SUBSTITUTED
OLEFINS, AS ADDITION-
SUBSTRATES; THE ORGANIC
CHEMISTRY NOTEBOOK, Nº 15**

**HIDROBORACIÓN, UNA BREVE
REVISIÓN HISTÓRICA A TRAVÉS DE
VISTAS MECANICISTAS, PARTE I:
OLEFINAS ALIFÁTICAS Y
AROMÁTICAS COMO SUSTRATOS
DE ADICIÓN; EL CUADERNO DE
QUÍMICA ORGÁNICA, Nº 15**

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Short review

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ABSTRACT

The Organic Chemistry Notebook Series, a Didactical Approach, is the series designed with educational purposes in the organic synthesis field. With the present paper we add to a total of fourteen contributions so far in the series.

This series of studies is designed to help students when getting started in the synthesis subject. The method of learning includes many fully and explicitly designed reactions step by step. The best manner to understand a synthesis is by means of graphical views, in this case, the ones proposed by the authors of this series, and, when in most of the cases they are accompanied by illustrative comments that describe the graphical mechanistic proposals and add some criteria deduced from the different mechanistic steps. During the past 14 chapters we have taken a series of reactions compiled by W. Carruthers in 'Some modern methods of organic synthesis', and we have proposed didactical and mechanistic views for them. Now we are boarding another important compiling source in the synthesis studies: 'Advanced Organic Chemistry, Part B: Reaction and Synthesis' by Francis A. Carey and



Richard J. Sundberg. This theme is included in the chapter "Electrophilic additions to carbon-carbon multiple bonds", and references therein.

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RESUMEN

La Serie de Cuadernos de Química Orgánica, un Enfoque Didáctico, es la serie diseñada con fines educativos en el campo de la síntesis orgánica. Con el presente trabajo agregamos un total de quince contribuciones hasta ahora en la serie.

Esta serie de estudios está diseñada para ayudar a los estudiantes cuando se inician en el tema de síntesis. El método de aprendizaje incluye muchas reacciones diseñadas de forma completa y explícita paso a paso. La mejor manera de comprender una síntesis es mediante vistas gráficas, en este caso, las propuestas por los autores de esta serie y, cuando en la mayoría de los casos, van acompañadas de comentarios ilustrativos que describen las propuestas gráficas mecanicistas y agregan algunos criterios deducidos de los diferentes pasos mecanicistas. Durante los 14 capítulos anteriores, hemos tomado una serie de reacciones compiladas por W. Carruthers en "Algunos métodos modernos de síntesis orgánica", y hemos propuesto puntos de vista didácticos y mecanicistas. Ahora abordamos otra fuente importante en los estudios de síntesis: "Química orgánica avanzada, Parte B: Reacción y síntesis" de Francis A. Carey y Richard J. Sundberg". Este tema se incluye en el capítulo "Adiciones electrofílicas a los enlaces múltiples carbono-carbono" en el texto mencionado y referencias incluidas.

INTRODUCTION

Due to lack of knowledge of classical mechanisms, students experiment emptiness when comprehending a determined synthesis found in the literature. Since a mechanistic proposal is naturally mandatory for a rational explanation of products emerging from a synthesis, we offer the present series theming on mechanistic approaches on several published syntheses. As academics we are committed with the didactics and we have designed a series of articles exposing mechanistic theoretical proposals, articles have a character of short review. The present contribution: Hydroboration, a brief historical review through mechanistic views, part I: alkyl- and aryl-substituted olefins, as addition-substrates, The Organic Chemistry Notebook, N° 15, that refers to regioselectivity considerations, is the fifteenth study in the series: "The Organic Chemistry Notebook Series, a Didactical Approach" [1-14].

REVIEW OF REACTIONS, MECHANISTIC THEORETICAL PROPOSALS, DISCUSSION

Hydroboration of olefins

Hydroboration is a good method in many synthesis reactions employing electron-rich compounds as substrate, namely alkenes and alkynes [15]. Boron as borane BH_3 is a Lewis acid with a free orbital suitable for rooming an electron pair. Borane is available as a dimer, B_2H_6 , with the boron atoms bridged by two of the six hydrogens between them [15]. Due to its Lewis acid nature (Figure 1), boron interacts with Lewis bases such as oxygen, nitrogen or sulphur in ethers, tertiary amines and sulfides, $R_2O^+-BH_3$, $R_3N^+-BH_3$, $R_2S^+-BH_3$ [15], see boron hybridation's changement when borane overlaps with Lewis bases (Figure 2).

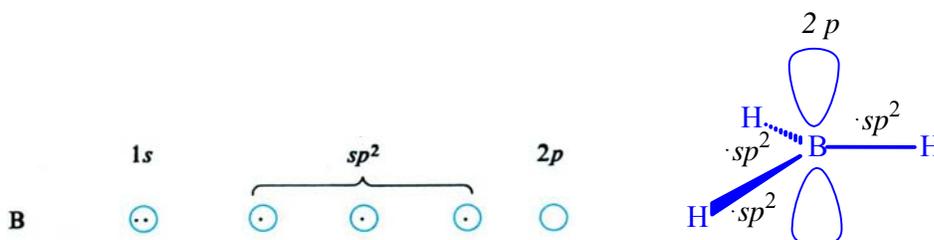


Figure 1. Boron electronic structure, borane structure

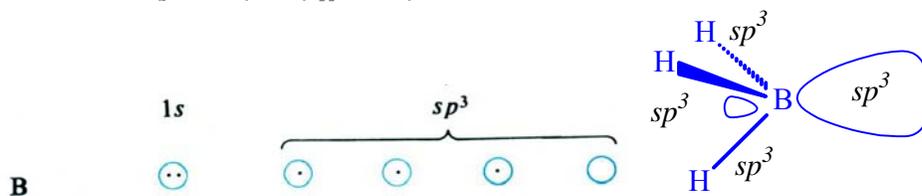


Figure 2. Boron hybridation's change when borane overlaps with Lewis bases

Using THF as solvent or DMS, the addition of borane over most alkenes is immediate [15], a reaction known as hydroboration. Hydroboration is stereospecific and very regioselective [15]. Steric and electronic effects are involved in the addition of borane and they favor the attack over the carbon the less substituted [15]. Since borane reacts with alkenes until completion of the hydride substitution in BH₃ and their replacement by the adding on three alkene molecules to give a trialkyl borane, the less substituted carbon is chosen by boron in order to avoid steric hindrance when adding the second and the third alkenic residue, *e.g.* the addition of BH₃ to 2-methyl-1-propene, where atom 1 is preferred over atom 2 for the new B-C bond [15], see Figure 3. This behavior gives this reaction its regioselective quality [15]. The minor option of addition of the boron atom to the carbon with the most steric hindrance represents severe nonbonded repulsions, see Figure 4.

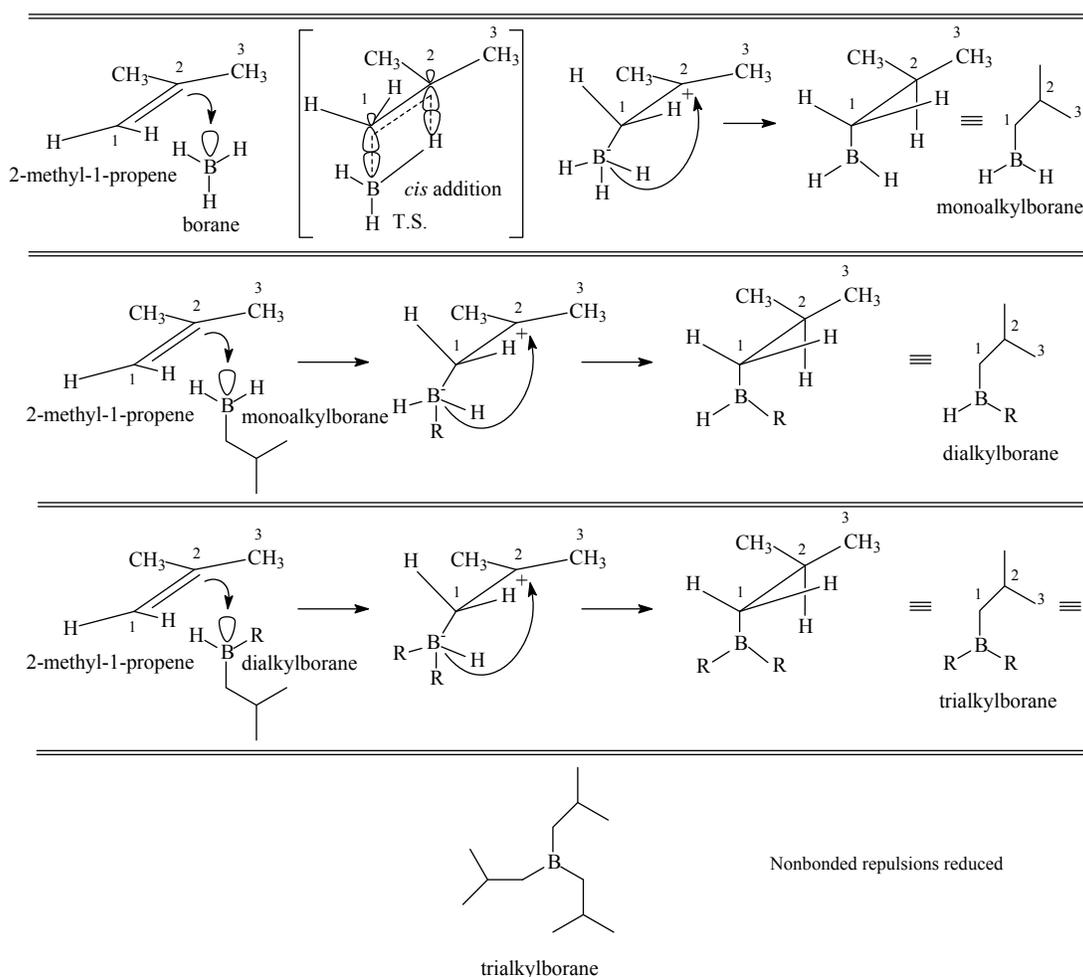


Figure 3. Hydroboration of 2-methyl-1-propene, boron's addition over the carbon with the less steric hindrance. Nonbonded repulsions reduced [15]. Theoretical mechanistic proposals by the authors

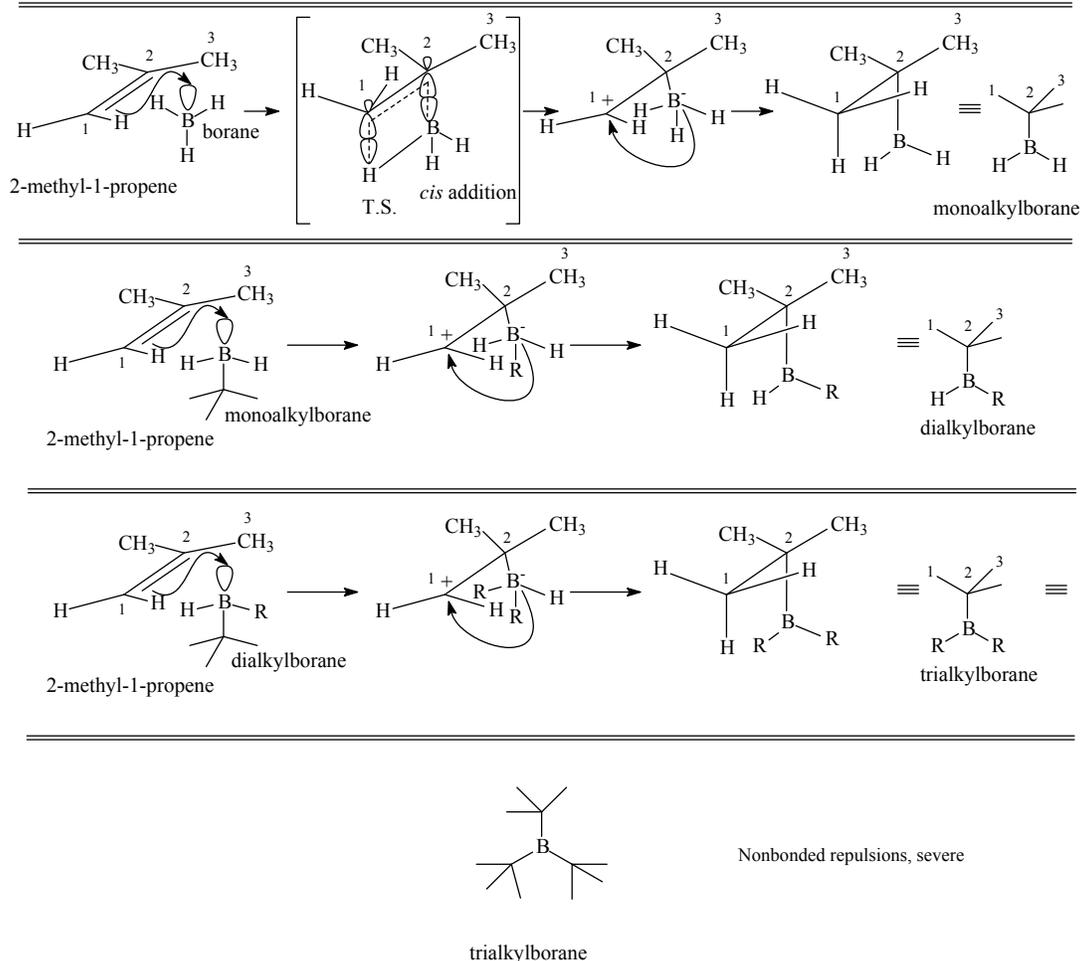


Figure 4. Minor hydroboration of 2-methyl-1-propene, boron's addition over the carbon with the most steric hindrance. Nonbonded repulsions, severe [15]. Theoretical mechanistic proposals by the authors

The regioselectivity can be improved by the use of some mono or dialkylboranes instead borane itself (disiamylborane [DAB, 15,16], thexylborane [TB, 15,17], 9-BBN [15,18]). While diborane shows a 94% of regioselectivity in hydroboration of 1-hexene, this increases to 99% when using chloroborane-dimethyl sulfide [15,19,20], or DAB [15,16] or thexylchloroborane-dimethyl sulfide [15,21], and 99.9% if 9-BBN is used [15,18], but it's equal (94%) if TB is used [15,17]. When the substrate of the hydroboration is 2-methyl-1-butene diborane gives 99%, chloroborane-dimethyl sulfide 99.5%, thexylchloroborane-dimethyl sulfide 97%, and 9-BBN 99.8%. Changing the substrate to 4-methyl-2-pentene diborane gave 57%, DAB 97%, TB 66%, thexylchloroborane-dimethyl sulfide 99%, and 9-BBN 99.8%. Finally, [15] styrene as substrate gave a regioselectivity of 80% (diborane), 98% (chloroborane-dimethyl sulfide, and DAB), 95% (TB), 99% (thexylchloroborane-dimethyl sulfide), and 98.5% (9-BBN). These alkylboranes can be prepared by hydroboration of the corresponding alkene, the degree of alkylation of borane comes out from stoichiometric control [15]. See Figure 5.

Comments

The fact of adding alkyl substituents on the boron atom implies obviously an increase of the steric hindrance at the moment of interaction between the reagent and the substrate olefin. This tool makes more difficult the interaction with the most hindered carbon and accelerates the rate of attacking the less hindered sterically carbon.

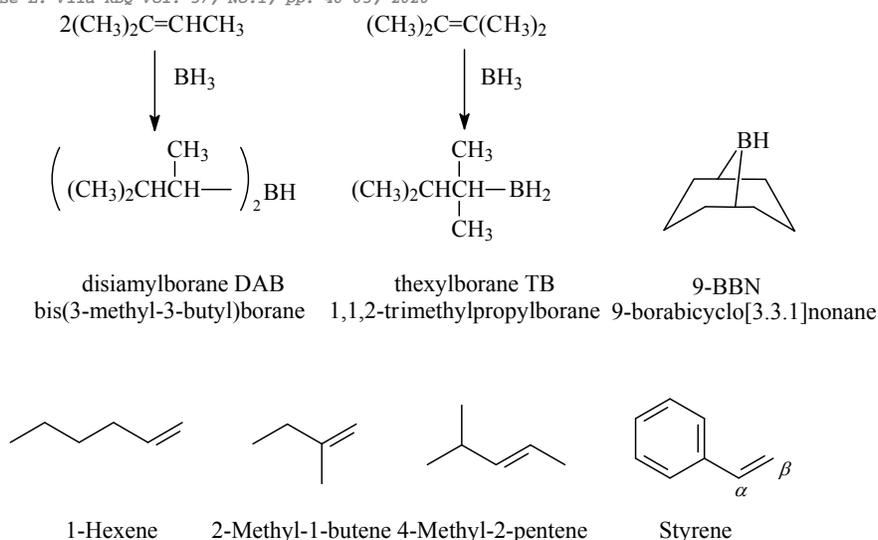


Figure 5. Regioselectivity, as shown by the use of many alkene substrates and varied borane alkyl derivatives [15,17]

As shown above in Figures 3 and 4, the borane or its derivative in hydroboration follows a *syn* (*cis*) stereospecific addition on the alkene [15]. The transition state (T.S.) involves a four center adduct with simultaneous bonding to B and H [15]. This atoms' spatial position makes possible the formation of the new bonds B-C and H-C, as being formed in the same side of the double bond [15]. A previous to the transition-state-step is the interaction of the two electrons of the bond π and the empty orbital p of the boron sp^2 , see Figure 6 [15,17].

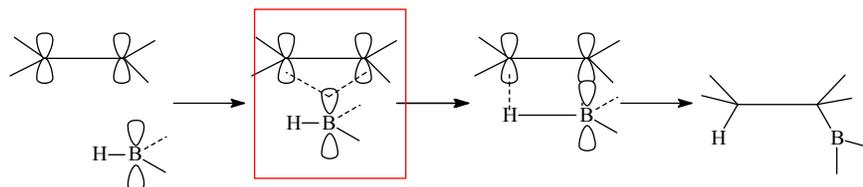


Figure 6. Previous to the transition-state-step is the *interaction of the two electrons of the bond π and the empty orbital p of the boron sp^2* [15]

Diborane is gaseous, but soluble in ether after dissociating into monomeric borane as etherate in varied ethers. Ether acts as catalyst in the addition reaction because of its ability of dissociating diborane [17]. The ethereal nature as solvent is not determining in the direction of addition in the hydroboration of styrene [17]. This means that the solvent does not participate actively in the transition state. On the other hand, the hydroborating agent (LiBH_4 or NaBH_4 or B_2H_6) does not make much difference in the direction of the addition either. Temperature variations didn't influence the addition direction [17]. As a general result in the distribution of boron over the double bond is from 94 to 98% in the same direction, or the less hindered carbon, in most of the cases. Indicated yields were of 90% or higher. The majority of olefins proceed in the reaction until the trialkyl borane, however, there are exceptions with the more hindered olefins that can reach the monoalkyl or dialkylborane stage [17]. The addition to styrene (20% α) or *p*-chlorostyrene (35% α) is different [17]. Styrene was chosen for studying the influence of temperature, solvent and hydroborating agent with regard to the direction of addition. A small increase in temperature produced an increase in the percentage of the α -derivative [17]. Hydroboration was also surveyed under the scope of the effect of the structure of the olefin involved in five groups, alkyl monosubstituted terminal olefins (I), alkyl disubstituted terminal olefins (II), alkyl disubstituted internal olefins (III), alkyl trisubstituted internal olefins (IV), *p*-substituted styrenes (V) [17]. The results pointed out the preferred anti-Markownikoff type of addition. Group I with alkyl substituents, linear or branched gave 93-94% of addition of the boron to the terminal carbon. If the substituent includes somehow phenyl, the boron addition to the terminal carbon is between 80 to 90%. Group II presents a 99% of the primary alkylborane due to two substituents that exert steric compression on the more unsaturated carbon. Group III like 2-



pentene and 2-hexene undergo addition of the boron in almost equal proportions to both sides of the double bond. Slight variation is seen from *cis* to *trans* isomers. Group IV shows addition results with boron added to the secondary carbon in about 98% and only 2% for the tertiary compound. Group V included results derived from the relative strong influence of the phenyl group over addition on the double bond on styrene.

Hydroboration of alkyl-substituted olefins

The directive effects of the similar structural features of different alkyl-substituted olefins during addition of hydrogen and boron to the double bond were studied [17].

Alkyl-monosubstituted terminal olefins $RCH=CH_2$

Straight-chain terminal olefins like 1-butene (Figure 7), 1-pentene and 1-hexene (Figure 8) give preferably the addition of boron to the terminal carbon in a 93-94% (6 to 7% for the secondary alkyl boron derivative) [17].

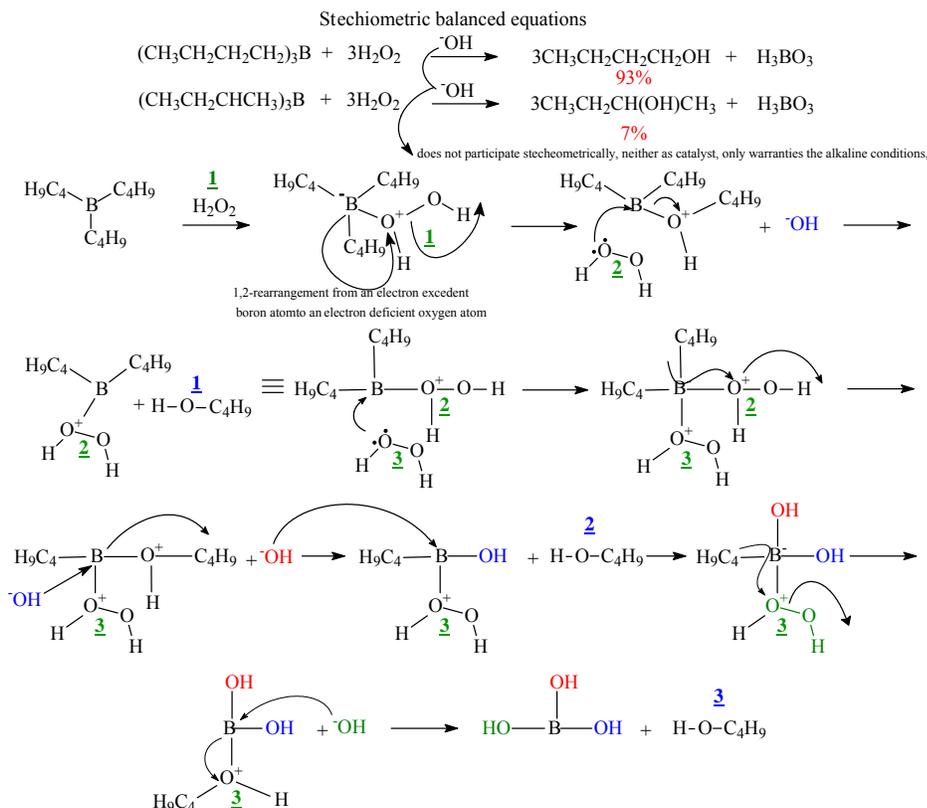
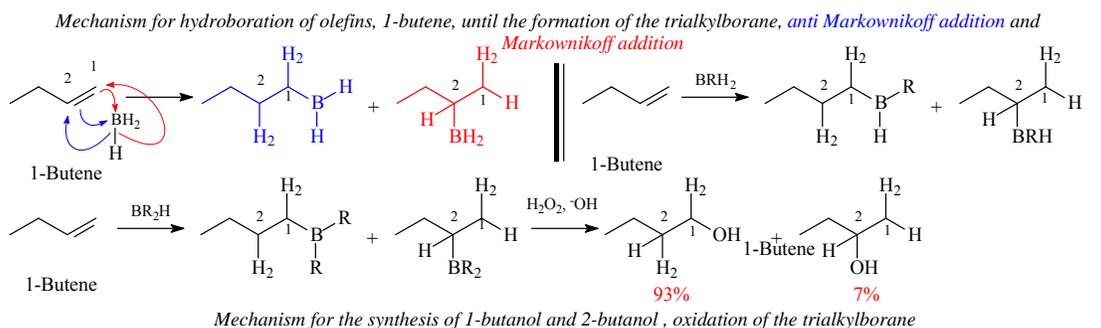


Figure 7. Hydroboration of 1-butene until completion of reaction obtaining the corresponding trialkylborane, oxidation until obtaining of the alcohol [17]. Theoretical mechanistic views proposals by the authors.



The alkyl branching makes not much difference in the boron distribution as shown with the use of 3-methyl-1-butene, 3,3-dimethyl-1-butene (Figure 9) and 4,4-dimethyl-1-pentene [17] (Figure 10).

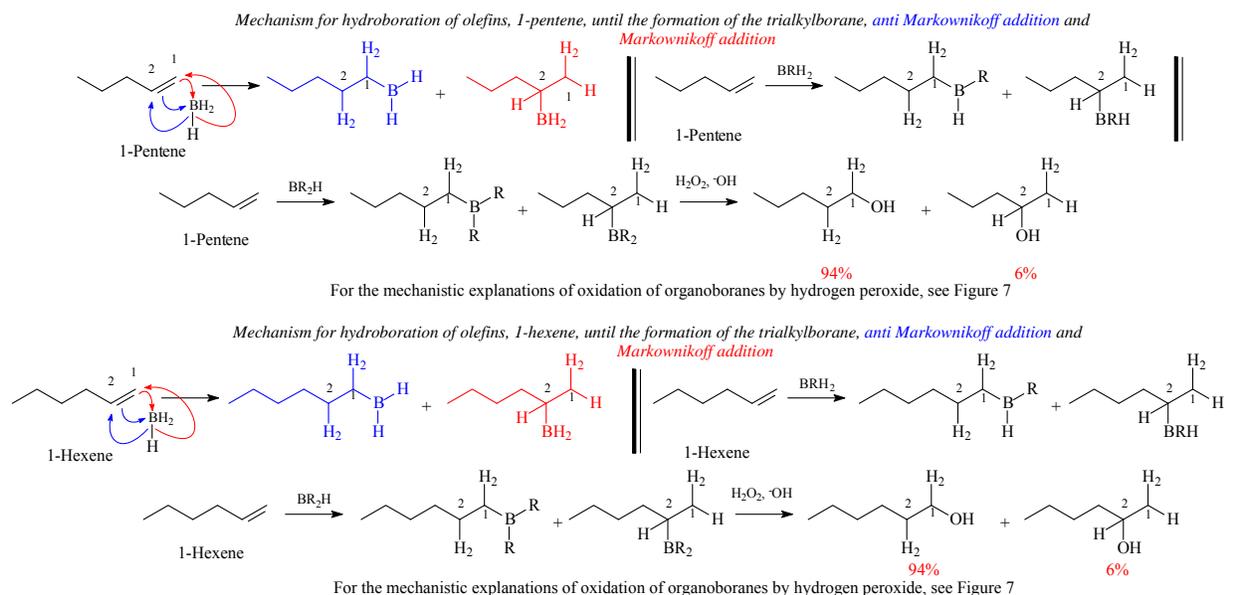


Figure 8. Hydroboration of 1-pentene and 1-hexene until completion of reaction obtaining the corresponding trialkylborane, oxidation until obtaining of the corresponding alcohol [17]. Theoretical mechanistic views proposals by the authors

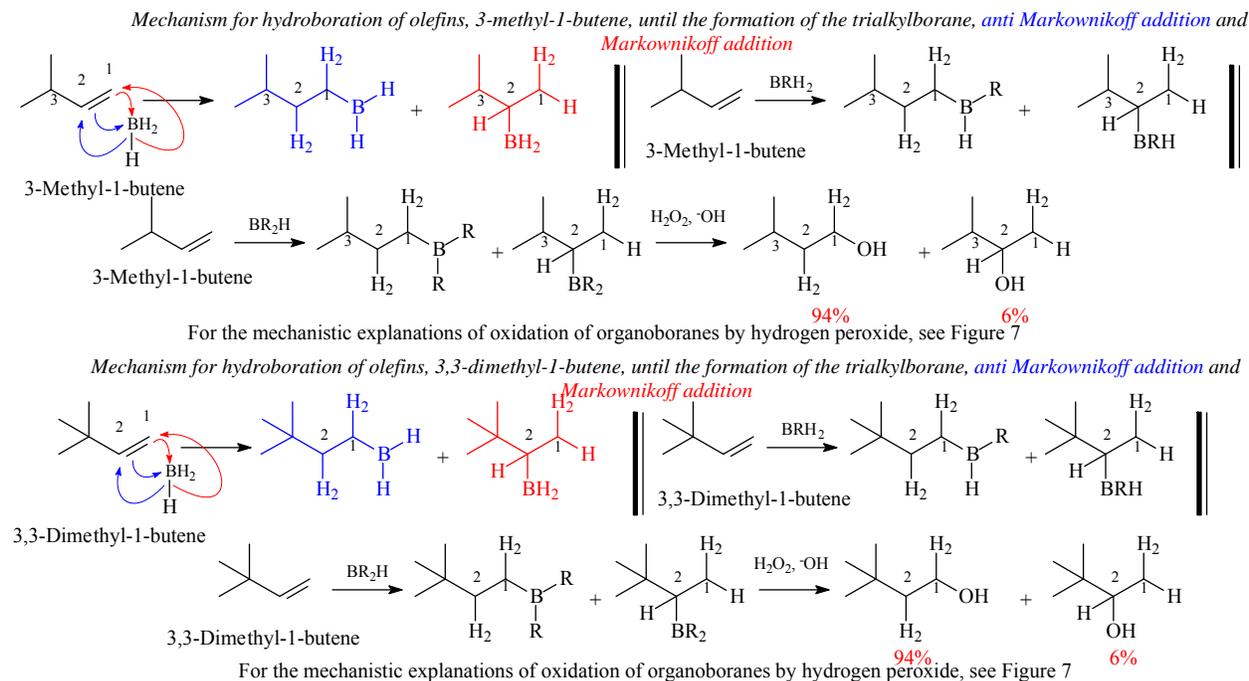


Figure 9. Hydroboration of 3-methyl-1-butene and 3,3-dimethyl-1-butene until completion of reaction obtaining the corresponding trialkylborane, oxidation until obtaining of the corresponding alcohol [17]. Theoretical mechanistic views proposals by the authors

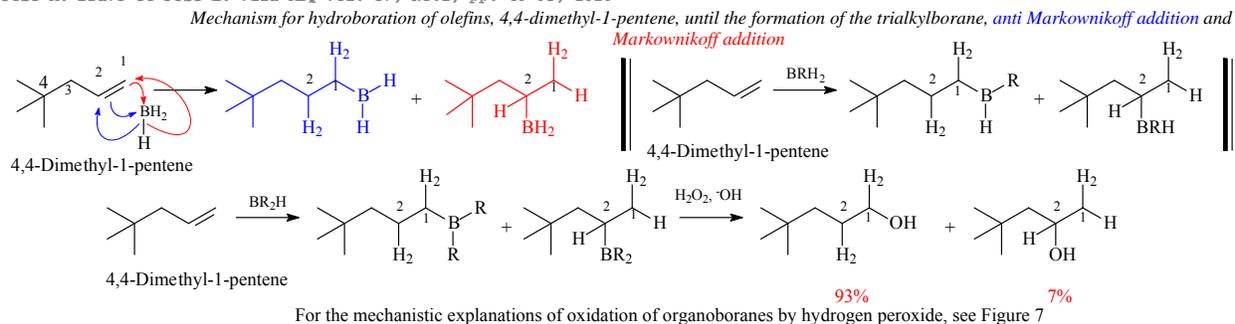


Figure 10. Hydroboration of 4,4-dimethyl-1-pentene until completion of reaction obtaining the corresponding trialkylborane, oxidation until obtaining of the corresponding alcohol [17]. Theoretical mechanistic views proposals by the authors

Alkyl-disubstituted terminal olefins $R_2C=CH_2$

Here we are faced to two geminal alkyl groups, very hindering, by the way. This provides a terminal methylene of alkenic carbon easily approachable by voluminous groups like borane and its derivatives. The directive effect of the geminal alkyl groups (identical or not), is overpowering giving rise to boron-addition results almost exclusively on the terminal carbon. As a sample let us mention that 2-methyl-1-butene and borane give 99% of the primary alkylborane and 1% of the tertiary alkylborane, and 2,4,4-trimethyl-1-pentene [17] (Figure 11).

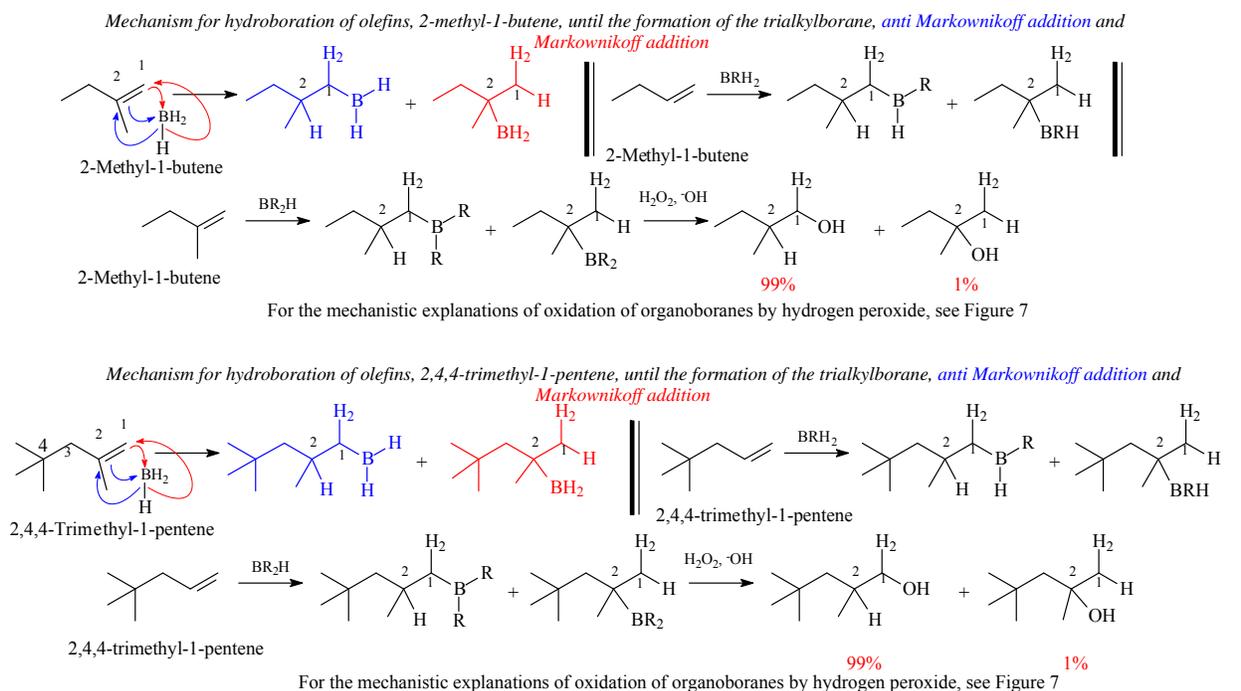
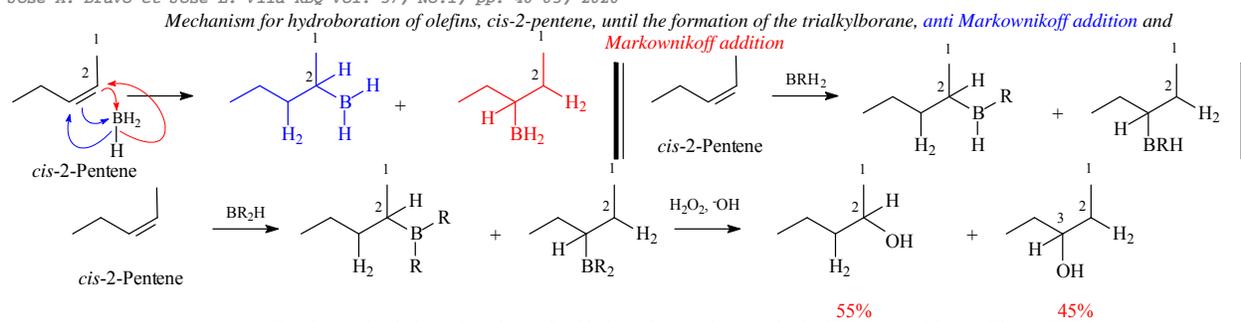


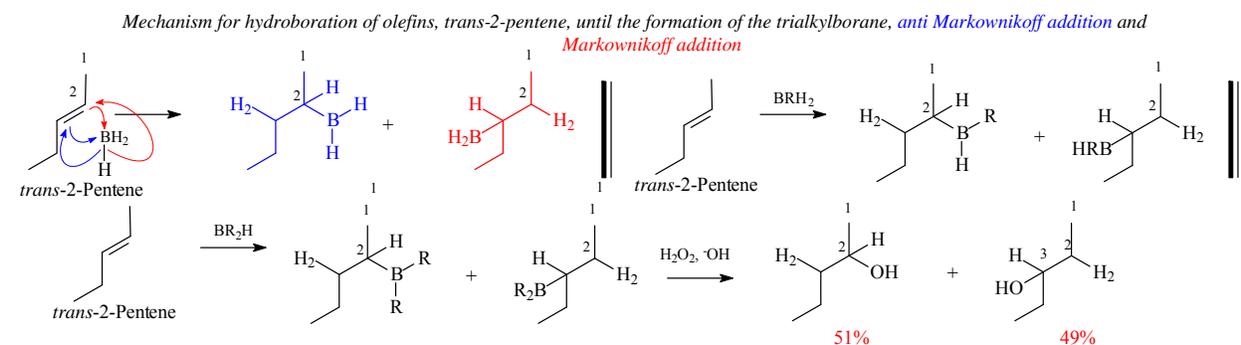
Figure 11. Hydroboration of 2-methyl-1-butene and 2,4,4-trimethyl-1-pentene until completion of reaction obtaining the corresponding trialkylborane, oxidation until obtaining of the corresponding alcohol [17]. Theoretical mechanistic views proposals by the authors

Alkyl-disubstituted internal olefins $RCH=CHR'$

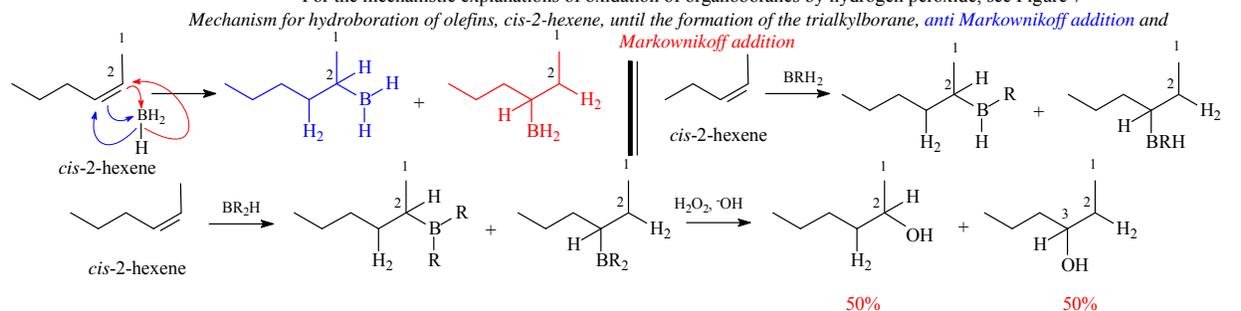
The immediate examples for the approach to alkyl-disubstituted internal olefins are 2-pentene and 2-hexene, which experiment hydroboration with boron addition fairly equally distributed on both extremes of the double C=C bond (Figures 12). The small difference in this distribution is observed in the geometrical isomers *cis* and *trans*.



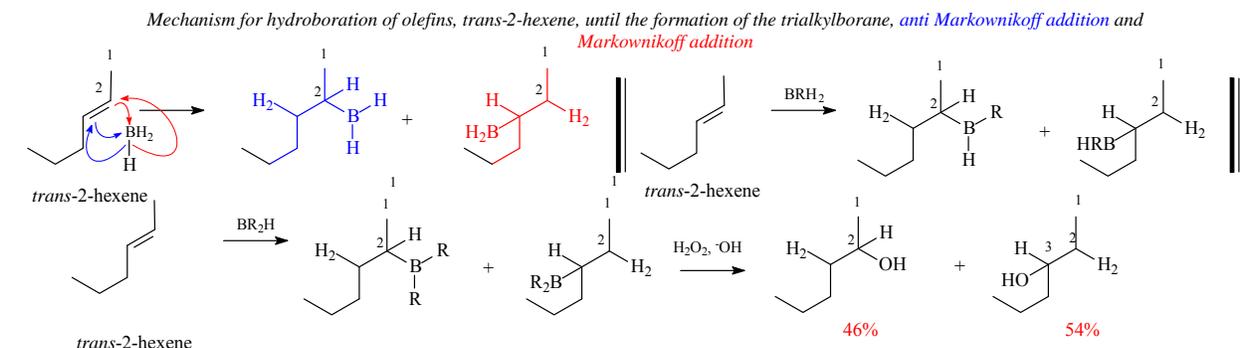
For the mechanistic explanations of oxidation of organoboranes by hydrogen peroxide, see Figure 7



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Figure 12. Hydroboration of *cis*-2-pentene and *trans*-2-pentene and *cis*-2-hexene and *trans*-2-hexene until completion of reaction obtaining the corresponding trialkylborane, oxidation until obtaining of the corresponding alcohol [17]. Theoretical mechanistic views proposals by the authors

Comments

A special case is the hydroboration-oxidation of *cis*-3-hexene [17], that gives a couple of enantiomeric alcohols after hydroboration and subsequent oxidation [17]. The distribution of boron on the double central C=C bond is



egalitarian (50:50). The result of hydroboration and hydrogen peroxide oxidation is the enantiomer couple: (*R*)-3-hexanol and (*S*)-3-hexanol, as a racemic mixture (50:50). The mechanism for the hydroboration-oxidation of *cis*-3-hexene is proposed in Figure 13. In this figure it is clearly established that, being *cis*-3-hexene a symmetric molecule, the alkyl side chains at each side of the double bond are identical. This means that there is no preference of the electrophile for either of the unsaturated carbons, giving a 1:1 distribution of borane. The article under current scrutiny and review [17] shows in Table IV the product of the hydroboration-oxidation of *cis*-3-hexene as an only product (100%) in the column of 3-ol result. This report is not specific, meaning by this that the only product is the 3-hexanol. However, Figure 13 clearly shows the formation of a racemic mixture. Figure 14 shows through molecular models the substrate structure, *cis*-3-hexene, and, the monoalkylboranes, product of the electrophilic addition, (*S*)-3-dihidroborexano (1) and (*R*)-3-dihidroborexano (2). Figure 15 shows the two enantiomers as specular images, demonstrating that the hydroboration gave two products instead of only one, as Table IV [17] seems to suggest.

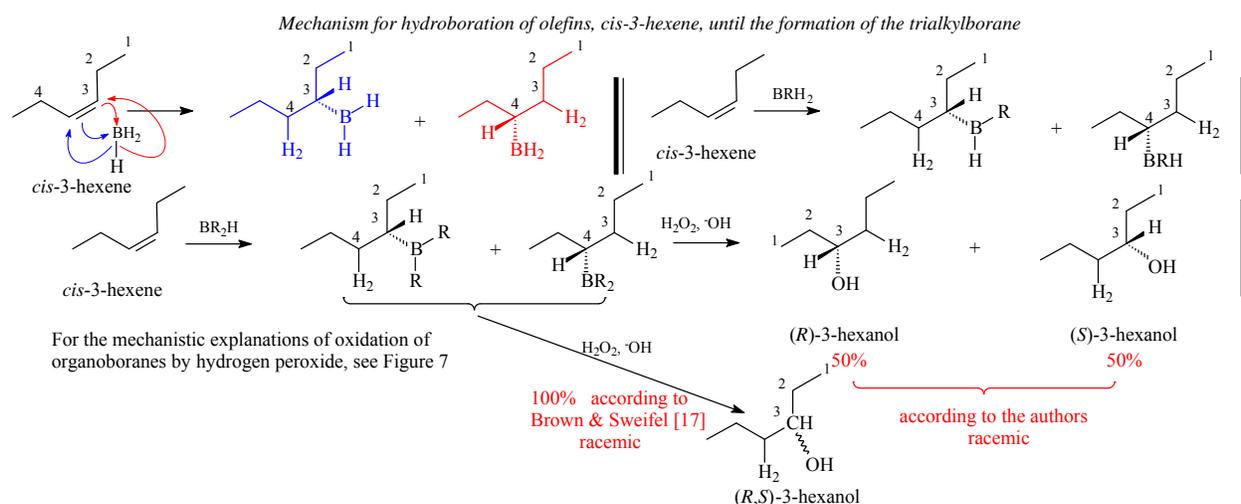


Figure 13. Hydroboration of *cis*-3-hexene until completion of reaction obtaining the corresponding trialkylborane, oxidation until obtaining of the corresponding alcohol [17]. Theoretical mechanistic views proposals by the authors

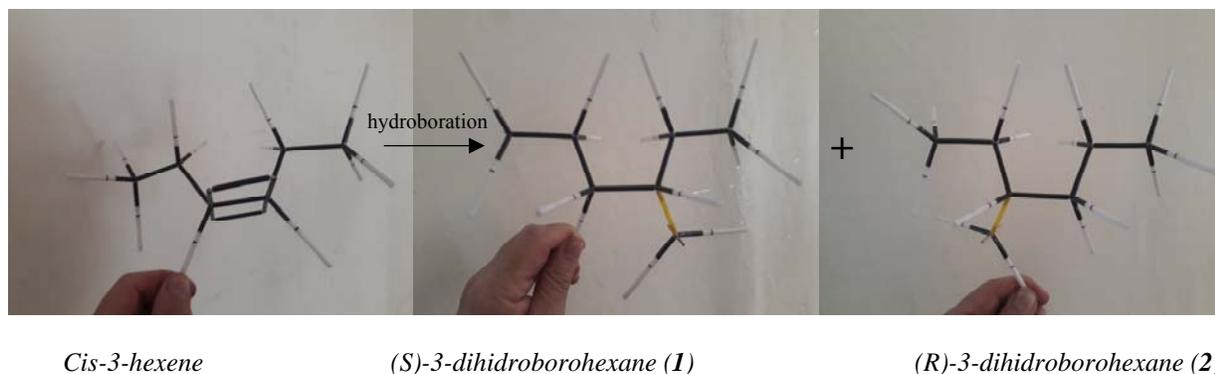


Figure 14. Hydroboration of *cis*-3-hexene to afford (*S*)-3-dihidroborexano (1) and (*R*)-3-dihidroborexano (2)

This analysis of the hydroboration of *cis*-3-hexene to afford a couple of enantiomers (*S*)-3-dihidroborexano (1) (*R*)-3-dihidroborexano (2) as a racemic mixture as shown in Figures 14 and 15 demonstrates existence of the transition state that directs a *syn* (*cis*) simultaneous addition of the H-B bond to the double bond. If the addition were *anti* (*trans*), it means in two instead of an only step, then the two addition-products or the 3-dihidroborexanos, would be diastereomers, and not enantiomers as it's the case.

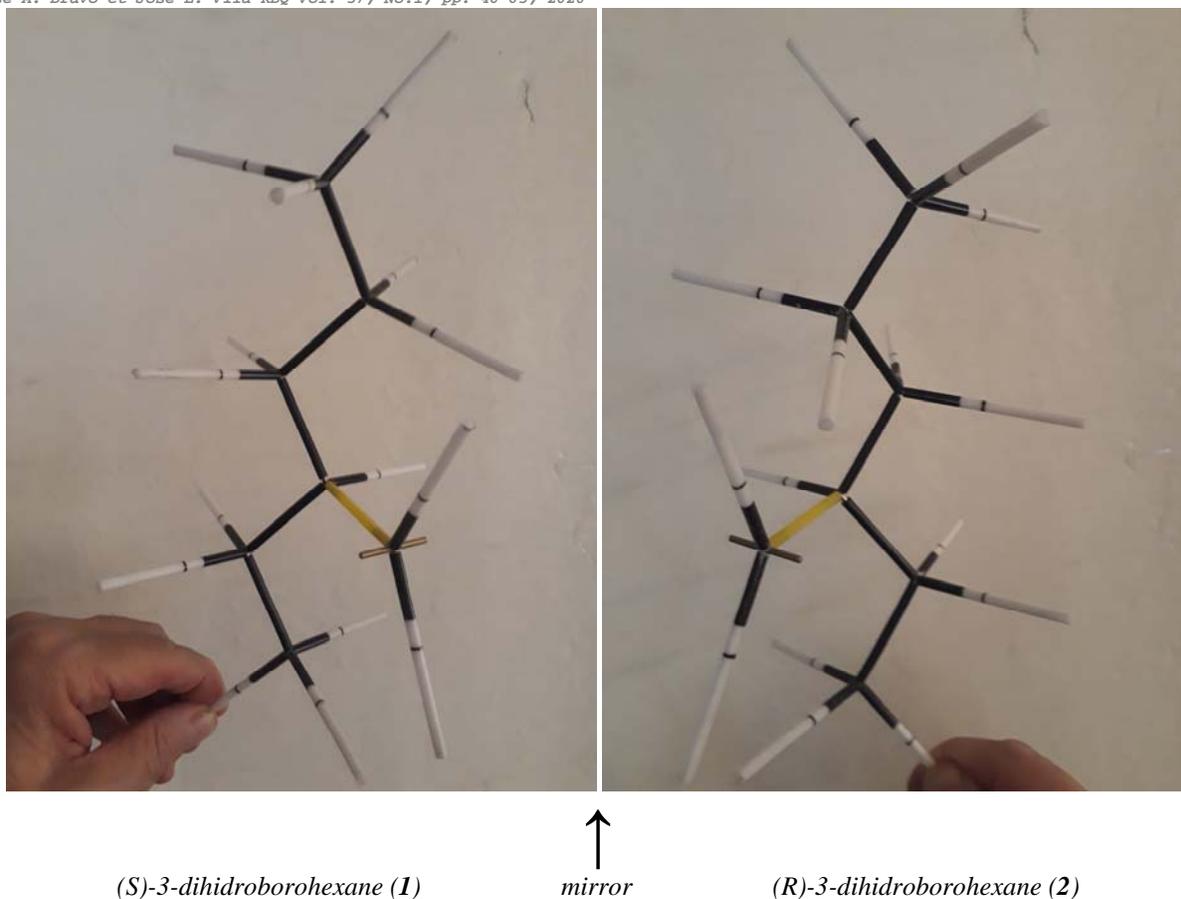


Figure 15. Enantiomeric products of hydroboration of *cis*-3-hexene, (*S*)-3-dihydroborohexane and (*R*)-3-dihydroborohexane as specular images

If the structural complexity at one of the side chains of the double bond is augmented, a small preference of the boron atom for the carbon atom with the less hindering side chain is manifested as shown for *trans*-4-methyl-2-pentene (Figure 16) and *trans*-4,4-dimethyl-2-pentene (Figure 17). The reaction proceeds until the dialkylborane in the case of *trans*-4,4-dimethyl-2-pentene [17].

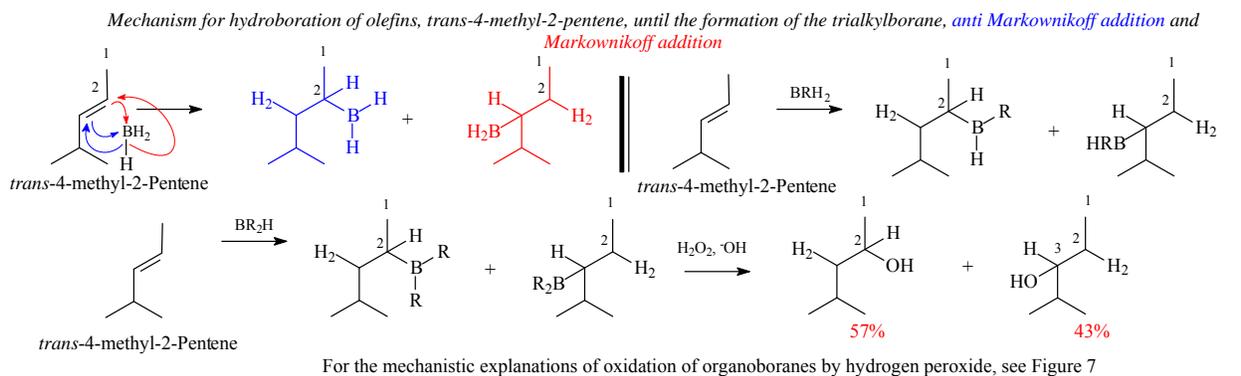


Figure 16. Hydroboration of *trans*-4-methyl-2-pentene until completion of reaction obtaining the corresponding trialkylborane, oxidation until obtaining of the corresponding alcohol [17]. Theoretical mechanistic views proposals by the authors

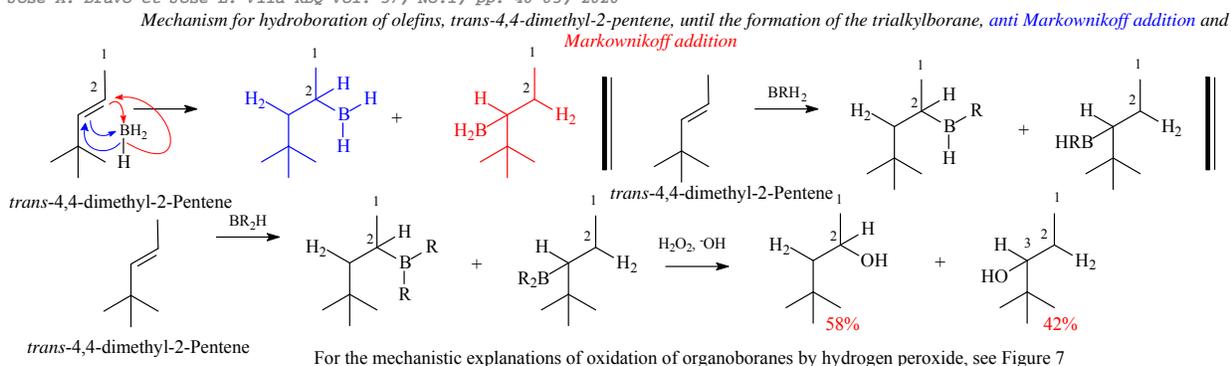


Figure 17. Hydroboration of *trans*-4,4-dimethyl-2-pentene until completion of reaction obtaining the corresponding dialkylborane, oxidation until obtaining of the corresponding alcohol [17]. Theoretical mechanistic views proposals by the authors

Alkyl-trisubstituted olefins $R_2CH=CHR'$

Two products were assayed as examples of hydroboration of alkyl-trisubstituted olefins, $R_2CH=CHR'$, 2-methyl-2-butene with a 98% of accommodation of borane on the secondary unsaturated carbon and 2% on the tertiary carbon, and 2,4,4-trimethyl-2-pentene, which afforded the same distribution on the secondary and tertiary carbon as the 2-methyl-2-butene. Both compounds under mild conditions reached until the second alkylation of borane only [17].

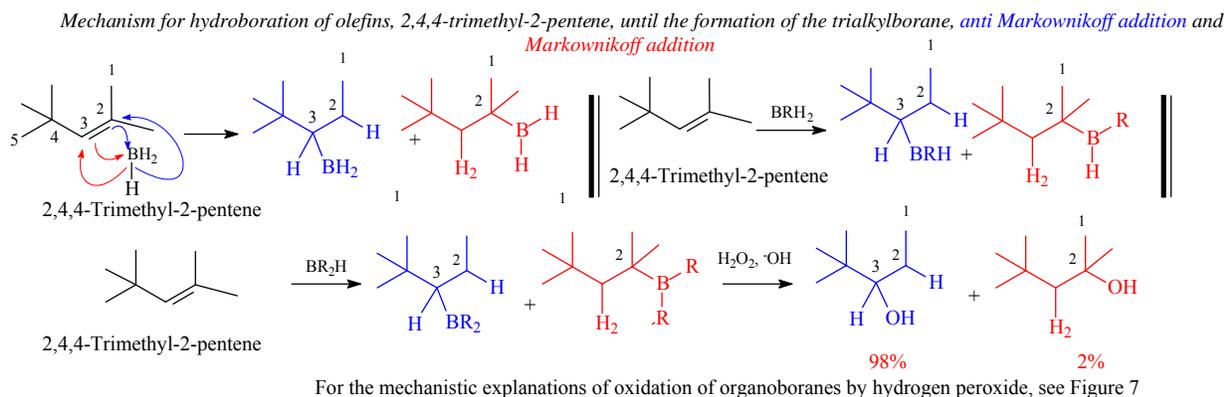
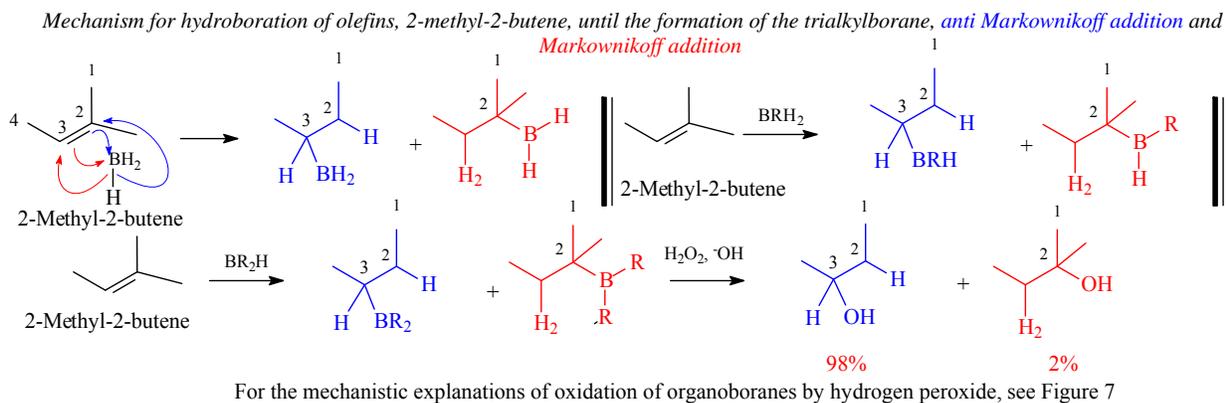


Figure 18. Hydroboration of alkyl-trisubstituted olefins: 2-methyl-2-butene and 2,4,4-trimethyl-2-pentene until completion of reaction obtaining the corresponding dialkylborane, oxidation until obtaining of the corresponding alcohols [17]. Theoretical mechanistic views proposals by the authors



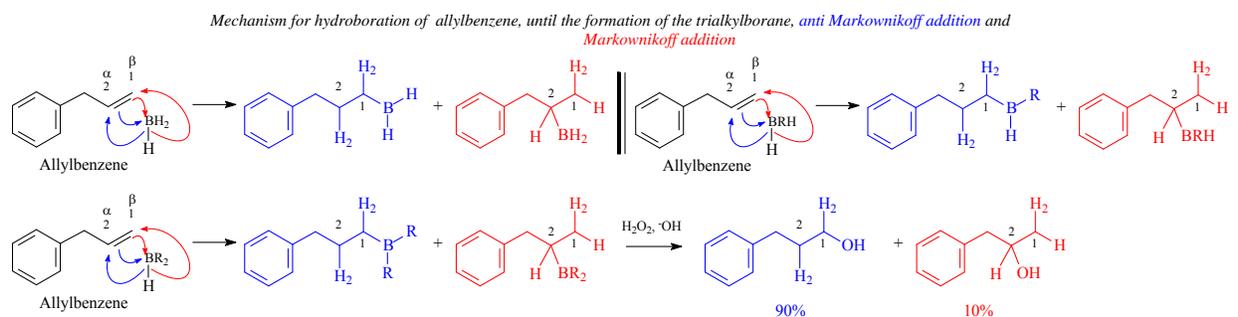
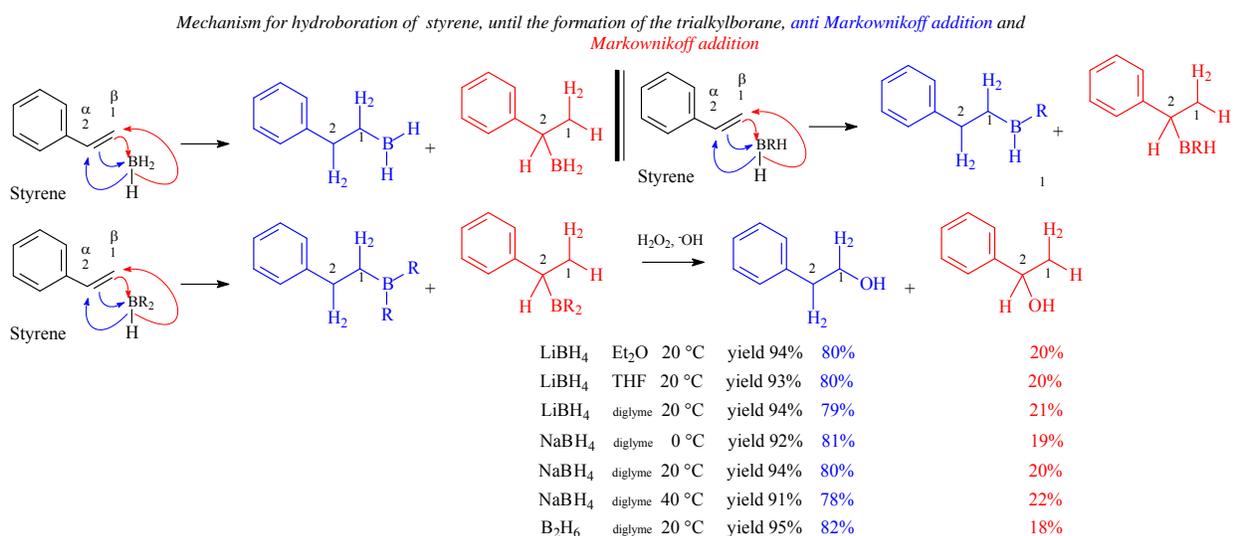
Hydroboration of aryl-substituted olefins

Influence of the hydroboration (reducing) agent employed and temperature in the hydroboration of styrene

The use of one or another hydroboration agent makes no remarkable difference in the addition distribution of the couple H-B on the double bond. Thence there was no significant change being noticed with the use of LiBH_4 (Et_2O , THF or diglyme), NaBH_4 (diglyme) or B_2H_6 (diglyme) [17]. A small increase in the proportion of addition on the α -carbon of styrene was produced with increasing temperature [17].

Styrene and its derivatives

Styrene gives much different addition results of the hydroboration agent than those of alkyl-substituted olefins with proportions about 20% and 80% on the double bond [17]. See Figure 19 for a comparison of products in different solvents and by using different hydroboration agents and at different temperatures [17].



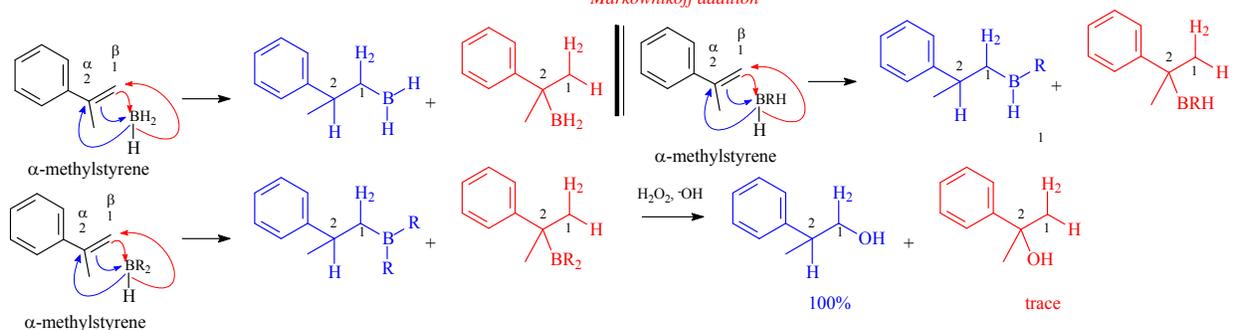
For the mechanistic explanations of oxidation of organoboranes by hydrogen peroxide, see Figure 7

Figure 19. Hydroboration of styrene and allylbenzene until completion of reaction obtaining the corresponding trialkylborane, oxidation until obtaining of the corresponding alcohols [17]. Theoretical mechanistic views proposals by the authors

As mentioned above, the distribution is 80% at C-1 and 20% at C-2 in the hydroboration of styrene. The effect of the phenyl group is minor but still present in allylbenzene with a 90_{C-1}:10_{C-2} distribution [17], see Figure 19. The case of α -methylstyrene shows the drastic directive effect giving a full regioselectivity with a 100% addition in position 1 [17]. See Figure 20.



Mechanism for hydroboration of α -methylstyrene, until the formation of the trialkylborane, *anti* Markownikoff addition and Markownikoff addition

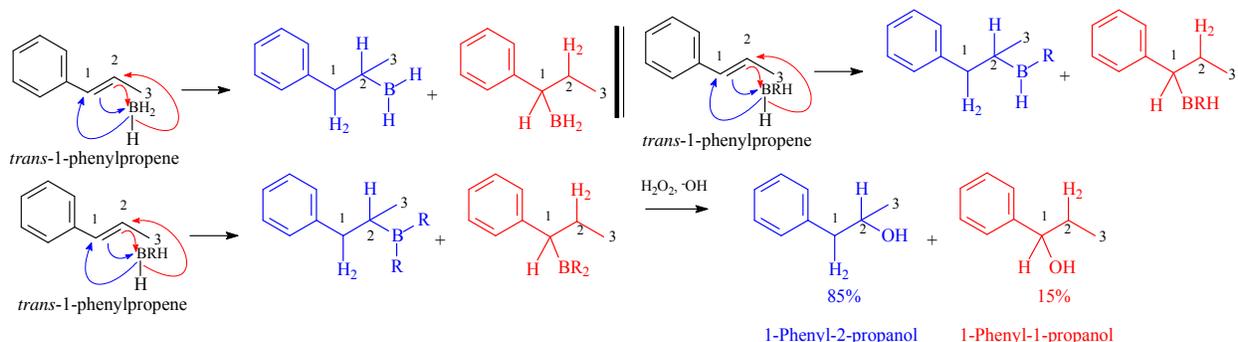


For the mechanistic explanations of oxidation of organoboranes by hydrogen peroxide, see Figure 7

Figure 20. Hydroboration of α -methylstyrene until completion of reaction obtaining the corresponding trialkylborane, oxidation until obtaining of the corresponding alcohols [17]. Full regioselectivity. Theoretical mechanistic views proposals by the authors

A comparison of the directive effect of alkyl-substituted (1-pentene) vs. aryl-substituted (styrene) olefins puts in evidence that the phenyl group of styrene is less effective than alkyl groups in directing the boron atom to the terminal carbon atom. This is seen in the hydroboration of *trans*-1-phenylpropene, with the competition between the methyl and the phenyl group [17], Figure 21.

Mechanism for hydroboration of *trans*-1-phenylpropene, until the formation of the trialkylborane, *anti* Markownikoff addition and Markownikoff addition



For the mechanistic explanations of oxidation of organoboranes by hydrogen peroxide, see Figure 7

Figure 21. Hydroboration of *trans*-1-phenylpropene until completion of reaction obtaining the corresponding trialkylborane, oxidation until obtaining of the corresponding alcohols [17]. Theoretical mechanistic views proposals by the authors

Comments

The article under current scrutiny [17], mentions a distribution of 15% for 1-phenyl-2-propanol and 85% of 1-phenyl-1-propanol, however, Figure 21 demonstrates that the distribution is exactly the opposite, namely 15% for 1-phenyl-1-propanol and 85% of 1-phenyl-2-propanol. This is an error of Brown and Zweifel [17].

p-Substituted styrenes

Due to the remarkable influence of the phenyl group in styrene on the tendency for the diborane when choosing a carbon atom over the other in the double bond, incited to the investigation of styrenes substituted in *para* [17]. This study was hoped to give enlightenment about the nature of the forces controlling the preference in the addition reaction.



The methyl substituent in styrene, provoked a minor effect, with a diminution from 20% in styrene in the α carbon to 18% in the *para* derivative. The methoxy substituent exerted a higher effect diminishing from 20% in the α position in styrene to 9% in the *para* derivative. Finally, the *p*-chloro substituent augments the distribution in the α position of the double bond until 35% from 20% in styrene [17], Figure 22.

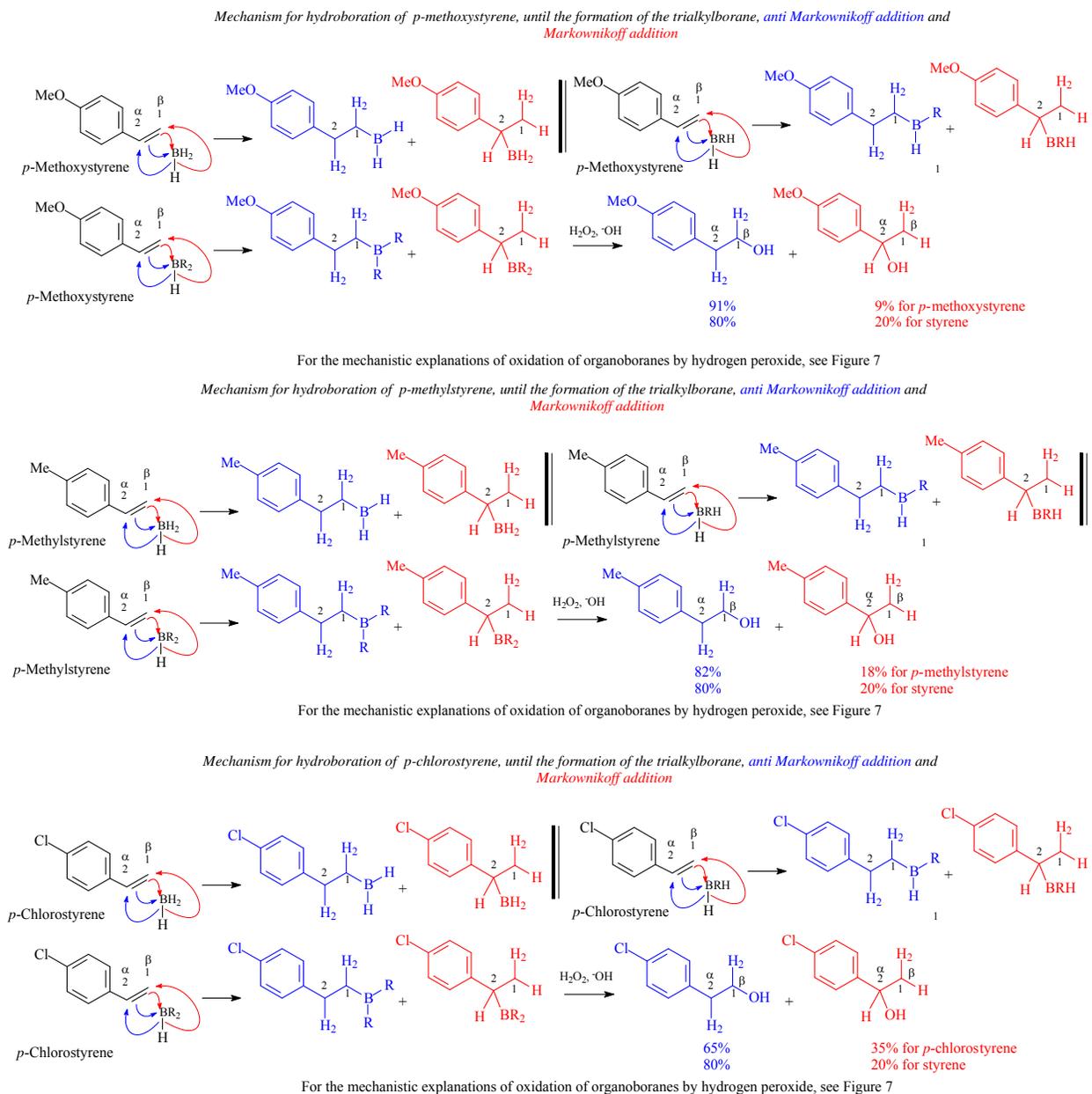


Figure 22. Hydroboration of *p*-methoxystyrene, *p*-methylstyrene and *p*-chlorostyrene in comparison to styrene, until completion of reaction obtaining the corresponding trialkylborane, oxidation until obtaining of the corresponding alcohols [17]. Theoretical mechanistic views proposals by the authors

CONCLUSION

The direction of addition of borane on olefins exposed in the examples above has been examined through the corresponding alcohols after oxidation [17]. The hydroboration repeats three times, each time with a new molecule



of substrate until reaching the trialkylated boron stage [17]. There is no evidence that the final distribution signaled by the alcohols is the same at each stage in the borane's alkylation pathway [17]. But, authors consider the distributions of the alcohols derived, as the same preferred by borane in the each of the preceding steps in the addition to olefins. Most of olefins proceeds to the trialkylborane R_3B [17]. Other olefins, in function of the alkyl increasing structural complexity arrive to the dialkylborane stage or the monoalkylborane stage [17]. All that have been exposed until here, regards with the directive effects derived from steric effects on the direction of addition of borane on olefins [17]. However, authors signaled that steric directive effects are not alone in exerting influence in the addition of borane on olefins [17]. Some of them being much different in branching, but with similar effect in distribution in results of the electrophilic addition, as seen in the following table [17].

RCH=CH₂	Terminal C	Internal C
R = ethyl	93%	7%
R = isopropyl	94%	6%
R = t-butyl	94%	6%
neopentylethylene	93%	7%
RCH=CHCH₃	R-terminal	CHCH₃
R = ethyl	51%	49%
R = isopropyl	57%	43%
R = t-butyl	58%	42%

All data collected and its previous analyses require an explanation other than the steric effects, and it concerns the electronic influences clearly playing an important role in controlling the direction of addition of the B-H bond to the double bond [17].

As shown above (figures 3 and 4), the addition of diborane to olefins happens clearly in a *cis* manner (see also figures 13-15), involving a four-center transition state [17]. The B-H bond is polarized with the hydrogen possessing a hydride character. Hence, by analogy, the electronic shifts are like in the addition of H-Cl to double bonds. See Figure 23, where borane adds to propylene, that kind of electronic shift is the explanation for the addition of boron to the terminal carbon. The same explanation is applicable to 2,2-dialkylethylenes (99% C-1) and for the secondary position in trisubstituted olefins (98%).

Propylene, a tetrahedral molecule, more-hindering

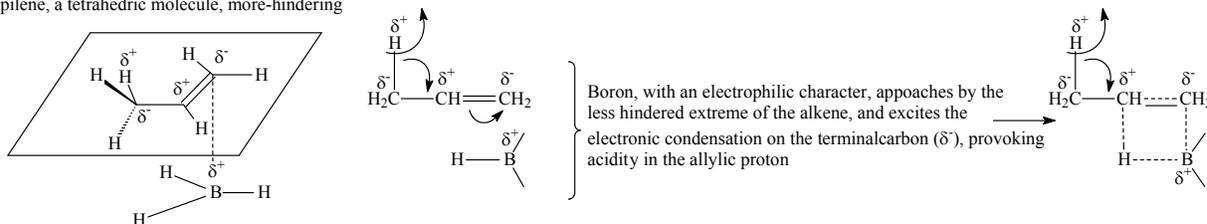
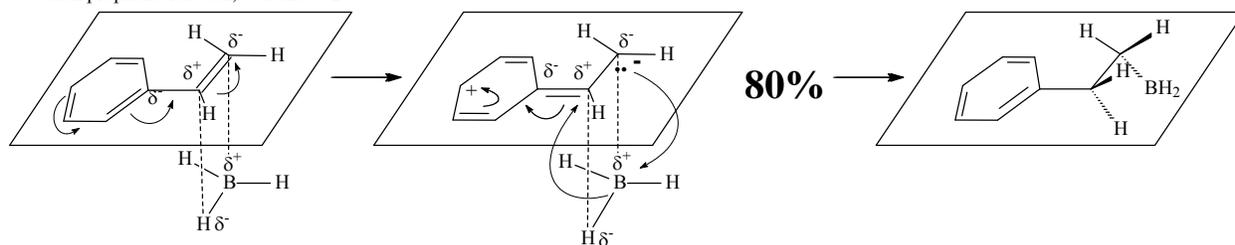


Figure 23. Hydroboration of propylene, electronic influences clearly playing an important role in controlling the direction of addition of the B-H bond to the double bond [17].

A similar approach can be employed for styrene to explain the preference of the boron atom for the ending position in the side chain. This approach also let explain the higher preference of C-2 with respect of the similar carbon in alkyl-substituted olefins: "In order to account for the enhanced substitution which occurs in the α -position, we must recognize that the phenyl group can also stabilize a negative charge in the α -position as in the benzyl anion. This transition state would be stabilized by an electron-withdrawing substituent, such as *p*-chloro-, and rendered less stable by an electron supplying substituent, such as *p*-methoxy" [17]. See Figure 24. The transition state at 80% possesses three resonant structures (two of them corresponding to benzene), in contrast to transition state at 20% that dispose of only 2 resonant structures (corresponding to benzene), being thus less stable.

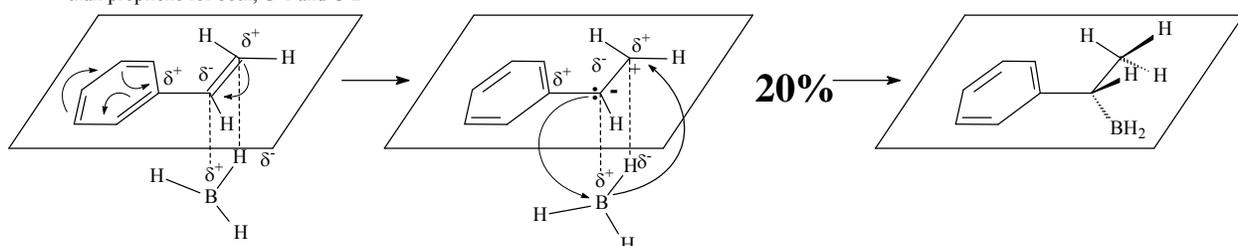


Styrene, a planar molecule, less-hindering than propylene for both, C-1 and C-2



80% approximation of borane (electrophile) to C-1 provoking electronic concentration on C-1 and on the C-1' of phenyl, and provoking an electronic deficit on C-2. The phenyl group, rich in electrons is able to stabilize the transition state as shown above

Styrene, a planar molecule, less-hindering than propylene for both, C-1 and C-2



20% approximation of borane (electrophile) to C-2 provoking electronic concentration on C-2 and an electronic deficit on C-1 and on C-1' of phenyl. The phenyl group, rich in electrons is able to stabilize the transition state as shown above

Figure 24. Hydroboration of styrene, electronic influences clearly playing an important role in controlling the direction of addition of the B-H bond to the double bond [17].

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